

Name: _____

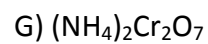
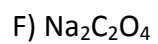
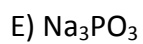
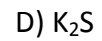
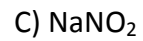
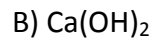
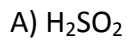
Oxidation and Reduction Reactions

1. On which side of an oxidation half reaction are the electrons?

2. On which side of a reduction half reaction are the electrons?

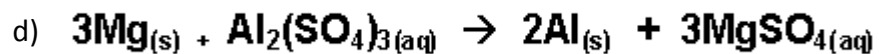
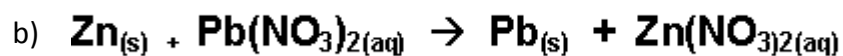
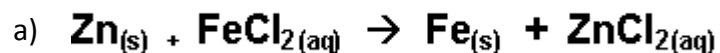
3. Explain why the oxidizing agent undergoes reduction in a redox reaction.

4. Determine the oxidation states on each element in the substances below.



5. For each of the following reactions:

- Write a balanced net ionic equation for the reaction.
- Identify which reactant is oxidized and which one is reduced.
- Identify which reactant is the oxidizing agent and which one is the reducing agent.
- Write balanced half reactions for the oxidation and reduction.



6. For each of the following reactions:

- Write the oxidation and reduction half reactions
- Write the overall redox equation
- Note: not all of these reactions are redox reactions

